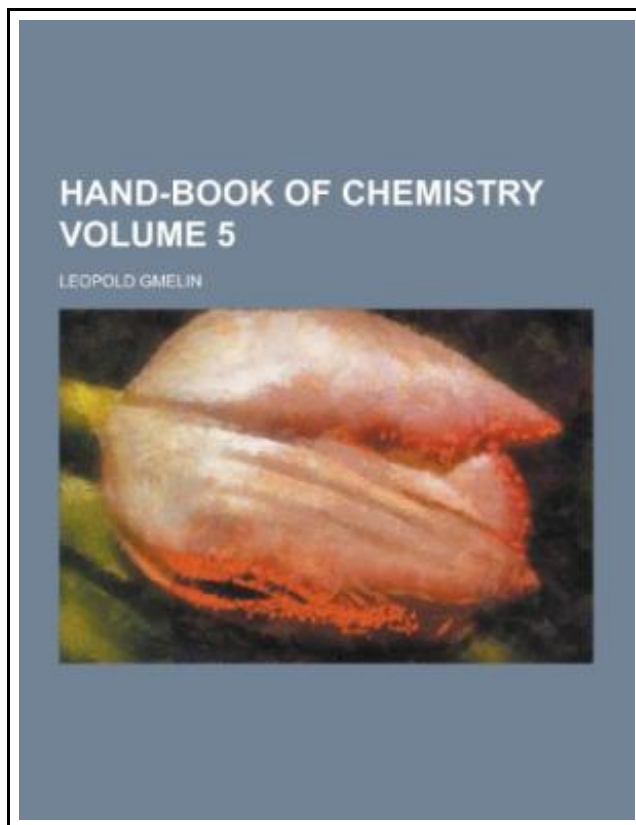


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(Kade Ankunding)

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Theclassics.Us, United States, 2013. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can usually download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1851 edition. Excerpt: .in a cast-iron crucibleand frequently renewed. (Gahn.)--4. A finely divided mixture of scaleoxide of iron and sulphur is heated to redness.--5. Iron pyrites is heated to redness, either alone or with a small quantity of iron-filings. The methods 2, 3, 4, and 5 may likewise yield the following sulphur compound, corresponding to magnetic pyrites; but when sulphide of iron is heated till it fuses, the residue, according to Bredbcrg, always consists of protosulphide of iron.--6. By igniting anhydrous ferrous sulphate in a crucible lined with charcoal. The product contains, however, only 34 per cent. of sulphur, likewise more or less carbon. (Berthier, Ann. Chim. Phys. 22, 241.)--7. By igniting 1 part of iron-filings with 1 part of charcoal-powder and 4 parts of load-sulphate, whereby metallic lead is obtained at the same time. (Fischer, N. Tr. 6, 1, 298.)--Yellow, with metallic lustre; yields a yellowish powder; it is attracted by the magnet. The protosulphide obtained by fusing iron pyrites is less hard than the bisulphide, and, on exposure to the air, crumbles to a bulky powder which emits a disagreeable odour when touched with wet hands. (Bredberg, Pogg. 17, 271.)--When iron pyrites is ignited in a current of hydrogen gas, the protosulphide is obtained in the form of a blackishgrey, non-magnetic powder, which dissolves in hydrochloric acid without separation of sulphur, and at the same time gives off pure hydrosulphuric acid gas. (H. Rose, Pogg. 5, 533.) The protosulphide obtained by igniting magnetic pyrites in...



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